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### **Chemical Engineering Kinetics**

POWER LAW BASICS | 6 2 POWER LAW BASICS 21 ARRHENIUS EQUATION The main assumption behind the Arrhenius expression is that  $k = A \cdot e^{-E_a/RT}$  This is an approximation, but it works quite well The rate coefficient is the term that is a function of temperature but may also depend

### **KINETICS Practice Problems and Solutions**

KINETICS Practice Problems and Solutions d Write the rate law for the overall reaction rate =  $k [A]^2 [B]^2$  9 Consider the following mechanism  $O_3 \rightarrow O_2 + O$  (fast)  $O_3 + O \rightarrow 2 O_2$  (slow) a Write the overall balanced chemical equation  $2 O_3 \rightarrow 3 O_2$  b Identify any intermediates within the mechanism c What is the order with respect

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